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SOLUTIONS MANUAL States of Matter Section 12.1
Gases pages 402–410 Practice Problems page 405 1.
Calculate the ratio of effusion rates for nitrogen (N₂)
and neon (Ne). $\frac{\text{Rate N}}{\text{Rate Ne}} = 0.849$ $\frac{\text{Rate N}}{\text{Rate Ne}} = \sqrt{\frac{\text{molar mass Ne}}{\text{molar mass N}_2}} = \sqrt{\frac{20.18}{28.02}}$
0.849 2. Calculate the ratio of diffusion rates for carbon
monoxide and carbon dioxide. 1.25 $\frac{\text{Rate CO}}{\text{Rate CO}_2} = \sqrt{\frac{\text{molar mass CO}_2}{\text{molar mass CO}}}$
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